Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_

Chapter 17 - Organization of the Periodic Table

1. The columns on the table are called groups or families, they share similiar properties.

2. The rows are called periods, the row number indicates the maximum number of energy levels an element has.

3. The table is numbered in order of increasing atomic number or number of protons.

4. Group 1 on the table has one valence electron (electron on the outermost energy level). This group will form a +1 cation when bonding.

5. Group 2 on the table has two valence electrons. This group will form a +2 cation when bonding.

6. Group 3 begins on the right hand side of the table and has three valence electrons. This group will form +3 cations.

7. Group 4 doesn't normally form ions, but if it does it can form a +4 or a -4 ion.

8. Group 5 forms -3 anions.

9. Group 6 forms -2 anions.

10. Group 7 forms -1 anions.

11. Group 8 elements do not form ions since they meet the octet rule, this means they have 8 valence electrons and are stable.